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Most adsorptive separation processes depend on differences in adsorption equilibrium. Kinetic separation processes are more difficult to design and require more precise control of the operating conditions, especially the cycle time. They are therefore generally employed only when no equilibrium selective adsorbent with sufficient selectivity and capacity is available. Nevertheless, several practically important separations are carried out using kinetically selective adsorbents and kinetic separation processes for the separation of C<sub>3</sub>H<sub>6</sub> / C<sub>3</sub>H<sub>8</sub> and N<sub>2</sub> / CH<sub>4</sub> appear to be on the verge of commercialization. A brief summary of these processes and the fundamental kinetic data on which they are based is given here.

### Olefin/Paraffin Separation

Recent studies of sorption kinetics in 8-ring zeolites (notably of the CHA family) have shown that subtle changes in the framework atoms (such as the substitution of Si for Al) have a small but significant effect on the dimensions of the 8-ring windows, leading to a dramatic effect on the diffusional activation energy and hence the sorption rate for critically sized molecules – see Table 1. It is clear that the diffusivity ratio ( $D_{\text{propene}}/D_{\text{propane}}$ ) for SiCHA and DDR3 is large enough to allow rapid uptake of the faster diffusing propene with very little adsorption of propane, thus providing the basis for a highly selective separation with high propene recovery.

**Table 1**

	<b>SiCHA</b>				<b>DDR3</b>			
	K	K <sub>ratio</sub>	D(cm <sup>2</sup> s <sup>-1</sup> )	D <sub>ratio</sub>	K	K <sub>ratio</sub>	D(cm <sup>2</sup> s <sup>-1</sup> )	D <sub>ratio</sub>
C <sub>3</sub> H <sub>6</sub>	700	0.8	8x10 <sup>-11</sup>	11,500	1000	~1.0	5.5x10 <sup>-12</sup>	9,500
	}		}		}			
C <sub>3</sub> H <sub>8</sub>	900		7x10 <sup>-15</sup>		~1000		6x10 <sup>-16</sup>	

### N<sub>2</sub> / CH<sub>4</sub> Separation

Efficient removal of N<sub>2</sub> from low grade natural gas requires a highly N<sub>2</sub> selective adsorbent. Most adsorbents adsorb N<sub>2</sub> and CH<sub>4</sub> at similar rates and with similar affinities but it has been shown that the titanosilicate ETS-4, when Sr exchanged and dehydrated at 540K provides the required kinetic selectivity. A pressure swing process based on this material has been demonstrated at pilot plant scale and appears economically attractive.

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